**CHEM 3A: Introductory General Chemistry**

**Fall 2014 Sections 55260**

**Lecture MWF (10-10:50am) in PHY 76**

**Lab Th (8:00-10:50am) in PHY 82 with R. Culp**

**Instructor:**      V. Cornel

**Contact info:**   e-mail veronica.cornel@reedleycollege.edu using “Chem3A” in subject line, or 638-3641 ext 3449

**Webpage:**       [**http://blackboard.reedleycollege.edu**](http://blackboard2.fresnocitycollege.edu/)

**Office Hours:**    PHY 78 (MW) 12-1pm and virtual office hour Fridays 9-10am

**Tutoring:** Free tutoring available in the Tutorial Center (by the library) or STEM tutorial center (FEM)

**Course Objectives**: Chemistry 3A is an elementary survey course in chemistry including lab work. It is designed to give the student a chemistry background for a wide variety of careers including forestry, nutrition, nursing, physical therapy, teaching and other biological and health related fields.

**Course Prerequisite: Math 103** Students will need to be familiar with basic algebra before taking this course as there is a lot of math involved.

**Course Advisories:** ENGL 1A, CHEM10 or high school chemistry

**Text and Materials:**

1.      Nivaldo J. Tro: “Introductory Chemistry” **3rd , 4th or 5th** Edition. The Mastering Chemistry CD is not needed.

2. Download fill-in notes and the labs weekly off Blackboard

You will need safety glasses ($5 at hardware store), and a calculator with “exp” (or “EE”) and “log” keys ($12 at Walmart), but not a programmable calculator or the pink STATS calculator.

**Lecture Notes:** The ability to listen carefully and to take good lecture notes in an essential college skill. Students should print out the fill-in notes, homework and lab assignments off my Blackboard website prior to coming to class. The more effort you put into your homework, the better you will do in exams.

**Homework:** Homework will be assigned every lecture. It is essential to your success in this class that you do all the assigned homework and read the relevant sections in your Textbook. All homework will be collected at the beginning of the following lecture and selected problems graded. This is to ensure that you work consistently and can apply what you learn to problems. There will be no make-up homework assignments, but I will drop the lowest four homework assignments. Do not just copy somebody else’s homework or you will not be able to do the problems for yourself in the exams. You can ask another student or tutor to help you start some problems, but you need to work them out for yourself. Even if you get all the problems wrong, you will still get 70% for the assignment for attempting all the problems yourself and showing all your work. You will learn where you are going wrong when I go over the homework. The latest I will accept homework is just before I hand back the graded homework the next lecture. This is not ideal as you won’t have your homework in front of you when I go over it and you will loose 10% for the homework being late. Absence is not an excuse for not doing your homework as you can send it in with another student, or count that assignment as one you drop. If you leave the class or are disruptive while I go over homework, I will also deduct points. It is advisable to write out the homework questions as well as the answers so you can study your homework. You can also do the corresponding odd number problems for extra practice and check the answers at the back of the book.

**Attendance:** Attendance in lecture and lab is mandatory. The student will be dropped automatically if she/he misses the first day of class, without contacting the instructor. If a student misses more than 25% of the lectures/labs, without contacting the instructor with a valid excuse, they will also be dropped. If you miss a lecture you need to read and summarize the chapter in the textbook **before** meeting with the instructor to discuss any problems. The homework will be on the internet notes so that you can do the homework even if you missed the lecture. **There will be no make-up exams**. The **final exam grade** will count as an exam and will also be counted for the grade for the missing exam. If you have not missed any exams, and do better in the final exam than one of the earlier exams, the final exam grade will **replace the earlier exam grade**. If you miss two exams you will receive a zero for the second missing exam. If a student is disruptive (including using cell-phones, interrupting the instructor continuously) they may be asked to leave the lecture/lab and recorded as "absent".

**Grading and Exams:**  There will be 5 exams covering the material of previous lectures and a final cumulative exam. These 6 exams will be equally weighted and count 65% all together (13% each).

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| Laboratory (25%): |  |
| Lecture Material (75%): | **Exams 65%** |
|  | **Homework 10%** |

The grading scale to be used is **A** 90-100%, **B** 80-89%, **C** 70-79%, **D** 60-69%, **F** 0-59%

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| **Student Learning Outcomes:** |
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| Upon completion of this course, students will be able to: |
| 1. Use dimensional analysis to solve for an unknown parameter of density, volume, mass, pressure, temperature, molar mass, concentration, or an empirical formula. 2. Construct and balance a chemical reaction and use the reaction to predict stoichiometric quantities. 3. Explain concepts from the periodic table and the use the periodic table to solve chemical problems. 4. Describe acid-base reactions and how to calculate pH. 5. Name and draw Lewis diagrams of inorganic and molecular compounds from the formula and vice versa. 6. Safely conduct laboratory experiments implementing concepts and principles learned in lecture. |

**Course Objectives**:

Upon completion of this course, students will be able to:

1. demonstrate an appreciation for the impact of chemistry on modern society and the relationship between chemistry and other disciplines including agriculture, the medical field, and industry;
2. classify types of matter, recognize physical properties and chemical properties, and a general understanding of the Law of Conservation of Mass and the Law of Conservation of Energy;
3. perform unit conversions using the correct significant figures; between the English and metric systems, temperatures in different units, density, energy, and with SI units;
4. use the periodic table to predict physical and chemical properties of elements and calculate molar masses of compounds and molecules;
5. recognize the electromagnetic spectrum and have a basic understanding of the quantum mechanical model of the atom;
6. demonstrate the ability to name inorganic compounds given their formulas, and write formulas given names;
7. distinguish and identify metals, non-metals, metalloids, and the elements of alkali metals, alkaline earth metals, halogens, noble gases, transition metals, and elements of the lanthanide and actinide;
8. distinguish and identify between different types of intramolecular and intermolecular forces of attraction present in various substances based on chemical formulas and structures;
9. write Lewis Electron-Dot Formulas and identify the shape using VSEPR method;
10. write and balance chemical equations, and use these equations along with stoichiometry and the mole concept to convert quantities (e.g. grams or moles) of a given substance into quantities of an unknown substance;
11. calculate, empirical formulas, and percentage composition given the appropriate data;
12. distinguish and balance chemical equations of different types of reactions;
13. perform calculations involving a limiting reagent and determining the percent yield;
14. predict the physical behavior of gases to pressure, temperature, and volume changes;
15. solve simple mathematical problems involving formula calculations related to gas laws;
16. use gas laws and stoichiometry to calculate quantities (e.g. moles, volume, grams) of gas produced or consumed during a reaction;
17. calculate various parameters of solutions including molarity, dilution techniques, percentage concentration, and density.
18. construct heating and cooling curves;
19. describe state and energy changes accompanying heating and cooling curves;
20. apply the principles of equilibrium in reversible reactions, saturated solutions, solutions of weak electrolytes and solutions of gases in solving related problems;
21. use solution properties and stoichiometry to calculate quantities (e.g. moles, volume, grams) produced or consumed during a reaction;
22. describe colligative properties of solutions (e.g. boiling point elevation, freezing point depression, and osmotic pressure);
23. define and identify acids and bases and be able to perform math calculations involving the pH;
24. determine the nature and applications for electron exchange reactions;
25. understand the structure of the atomic nucleus;
26. understand the fundamental types of nuclear radiation and the effects they have on biological systems
27. and demonstrate laboratory skills which include operating an analytical balance; calibrating and/or use fundamental lab equipment such as a thermometer, barometer, buret, pipette; recognizing use and limitations of laboratory glassware; recording and reporting observations; using error analysis techniques to evaluate certainty of data; use safety precautions and general laboratory procedures.

Please be aware of the following rules:

* Tardiness, cell-phone use, leaving early, and sleeping during lecture is considered disruptive behavior and will result in a partial or full absence being recorded. Students will need to sign the sign-in sheet within the first 10 minutes of class.
* Fraudulent behavior during exams is graded with a (0) zero.
* No homework may be handed in after I have returned it or gone over it in class.  No alternative homework will be given. I will drop the lowest two homework assignments though.
* No extra credit will be given. You need to work consistently from the beginning.
* Please turn your cell phones onto “silent buzzer” mode during lectures so as not to disturb the class. No cell phones or i-pods will be allowed during exams.

**If you have a verified need for an academic accommodation or materials in alternate media (i.e., Braille, large print, electronic text, etc.) per the Americans with Disabilities Act (ADA) or Section 504 of the Rehabilitation Act, please contact the Disabled Student Services as soon as possible.**

**CHEM 3A Spring 2015 - Cornel/Culp**

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| |  |  |  |  | | --- | --- | --- | --- | | Week | **Dates** | **Lectures** | **Thursday Lab** | | 1 | Jan 12-16 | Intro, Periodic Table  2.2-3 Scientific Notation & Significant Figures  2.6 Dimensional Analysis and Measurements | Safety, Safety Quiz.  Study Guide 1: Significant Figures | | 2 | Jan 19-23 | **Mon: Martin Luther King Day**  3. Matter, Physical and Chemical Changes  4. Atoms, Elements and Ions | Exp 1: The Mass of a Penny  Make-up safety lecture | | 3 | Jan 26-30 | 5. Chemical Nomenclature: Compounds  **Wed: Exam 1**  5. Chemical Nomenclature: Molecules | Exp 2: Mass Volume and Density. Spreadsheet for printing your Exp 2 graph is posted. | |  | Jan 30 | **Last Day to drop class to avoid a “W”** | | | 4 | Feb 2-6 | **Monday: Labor Day, no class**  5.11 and 6.1-6.1 The Mole  6.7-8 Percent Comp. and Empirical Formulas | Exp. 3 Specific Heat Capacity of an Unknown Metal | | 5 | Feb 9-13 | 5. Polyatomic Ions and Hydrates  7.1-7.4 Balancing Chemical Reactions  **Fri: Lincoln Day** | Exp 4: Relative Masses of Cu and Zn. | | 6 | Feb 16-20 | **Mon: Washington Day**  **Wed: Exam 2**  Percent Water in Hydrates | **Lab Quiz 1 (Safety, equipment, Exp 1, 2, 3 and**  **Study Guide 1)** Study Guide 3: Chemical Nomenclature (Assigned) | | 7 | Feb 23-27 | 7.9-7.10  Types of Reactions  7.7 Net Ionic Equations, Electrolytes  8. Stoichiometry | Exp 5: Determination of an Empirical Formula. Due at end of lab. | | 8 | Mar 2-6 | 8.6 Limiting Reactants and 8.7 Percent Yield  8. Thermochemistry  8. Calorimetry | Exp 13: Percent water in hydrates | | 9 | Mar 9-13 | 9. Electronic configuration  **Wed: Exam 3**  9.7, 9.9 and 10.2 Periodic Table Trends | Exp 6: Double Replacement Reactions. | |  | **Mar 13** | **Last Day to drop class with a “W” (letter grades assigned after this date)** | | | 10 | Mar 16-20 | 10. Bonding  10.5 Lewis diagrams  10.7 Geometry | Exp 7: Alum production from scrap aluminum.  Will be completed next lab | | 11 | Mar 23-27 | 11. Gases  11. Combined Gas law  11. Ideal Gas Law | **Lab Quiz 2** **(expt. 4, 5, 6, 13 and Study Guide 3)**  Study Guide 6: and Molecular Models: Lewis Structures VSEPR and Molecular Shapes. | |  | Mar30-Apr3 | **Spring Break** |  | | 12 | Apr 7-10 | 11. Gas Stoichiometry and Partial Pressure  **Wed: Exam 4**  13. Solutions, Dilutions | Exp 9: Molar mass of a volatile liquid. | | 13 | Apr 13-17 | 14. Acids and Bases  14.6 Titrations  Electronegativity, Polar Bonds and Molecules | Exp 8: Determination of Total Water Hardness | | 14 | Apr 20-24 | 12. Liquids, Solids and Intermolecular Forces 14. pH  14. pOH | Exp 11: Determination of the Heat of fusion for Ice. | | 15 | Apr 27-  May 1 | 15. Chemical Equilibrium  17. Radioactivity and Nuclear Chemistry | Expt. 10: Due today Lab Practical: Titration of Citric Acid in Mountain Dew. Turn in before leaving! | | 16 | May 4-8 | Review  **Wed: Exam 5** | Expt. 12: Spectroscopy of a Dye, The Chemistry of Color (Using a Spec 20) | | 17 | May 11-15 | Review | **Lab Quiz 3 (Exp 7, 8, 9, 10, 11 and Study Guide 6)** | | **18** | **May 20** | **Wed: Final Exam 10-11:50am (Cumulative)** | No lab | |
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