**SYLLABUS FOR CHEMISTRY 1A – SPRING 2014**

Section 52670

Lectures: MWF 11-11:50 (CCI 204)

Labs: T and Th 11-1:50 (PHY 82)

**Instructor**:       **Veronica Cornel**

**Contact info**:   e-mail  **vmcornel3@verizon.net** (using “Chem1A” as the subject or I will delete it) or leave message at the front desk or on my voicemail (559) 638-3641 ext. 3449

**Website**: http://blackboard.reedleycollege.edu

**Office Hours**:  9:30-10:30 MWF in PHY78

**Tutoring:** Free tutoring available in the Tutorial Center (by the library) or the STEM tutorial center (FEM) .

**Course Objectives**: Chemistry 1A is an general course in chemistry designed not only for chemistry majors, but also for biology, physics, chemical engineering, pre-medical and pre-pharmacy majors.  As a prerequisite students need to have passed **CHEM10, or High School chemistry**, with at least a C grade as well as basic algebra (**Math 103**).

**Textbook**:          Nivaldo J. Tro: Chemistry: A Molecular Approach (1st, 2nd **or** 3rd Edition). The first half of the book will be used in CHEM1A and the second half in CHEM1B. The Mastering Chem CD is not required for CHEM1A homework, but will be required for CHEM1B (3rd edition). You can purchase the CD separately next semester straight from the publishers.

**Lecture Notes:**  Fill-in notes will be posted on Blackboard the Friday before each week. Print the notes and bring them to class. The homework problems numbers are listed at the end of each day's notes in three separate sets for the three editions of the textbook. For the first week the actual problems will be written out in case students have not purchased their textbook yet.

**Lab Manual**:      The labs have been photocopied for you and are available as a set at the bookstore almost at the cost of photocopying.

**Other Supplies**:  A calculator is required (needs exponents and logs, but not a programmable calculator as the programmable ones will not be allowed in the exams). You will be able to borrow safety glasses and a lab coat for the semester if you don't have your own, and will need to wear closed shoes.

**Homework**: Homework will be assigned every lecture. It is essential to your success in this class that you do all the assigned homework and read the relevant sections in your Textbook. All homework will be collected at the beginning of the following lecture and selected problems graded. This is to ensure that you work consistently and can apply what you learn to problems. There will be no make-up homework assignments, but I will drop the lowest two homework assignments. Do not just copy somebody else’s homework or you will not be able to do the problems for yourself in the exams. You can ask another student or tutor to help you start some problems, but you need to work them out for yourself. Even if you get all the problems wrong, you will still get 70% for the assignment for attempting all the problems yourself and showing all your work. You will learn where you are going wrong when I go over the homework. The latest I will accept homework is just before I hand back the graded homework the next lecture. This is not ideal as you won’t have your homework in front of you when I go over it and you will loose 10% for the homework being late. Absence is not an excuse for not doing your homework as you can send it in with another student, or count that assignment as one you drop. If you leave the class or are disruptive while I go over homework, I will also deduct points. It is advisable to write out the homework questions as well as the answers so you can study your homework. You can also do the corresponding odd number problems for extra practice and check the answers at the back of the book.

**Attendance**: Attendance in lecture and lab is mandatory. As an incentive to attend lectures, an additional two homework assignments will be dropped at the end of the semester if a student attends 90% of the lectures. The student will be dropped automatically if she/he misses the first day of class, without contacting the instructor. If a student misses more than 25% of the lectures/labs, without contacting the instructor with a valid excuse, they will also be dropped. If you miss a lecture you need to read and summarize the chapter in the textbook **before** meeting with the instructor to discuss any problems. The homework will be on the internet notes so that you can do the homework even if you missed the lecture. There will be no make-up exams. The final exam grade will be counted for the grade for one missing exam. If you miss a second exam you will get a zero for that exam. If you have not missed any exams, and do better in the final exam than one of the earlier exams, the final exam grade will replace the earlier exam grade. If a student is disruptive (including using cell-phones, talking, interrupting the instructor continuously) they may be asked to leave the lecture/lab and recorded as "absent".

**Drop Date to avoid a “W”:** The last day to drop this class to avoid a “W” is Jan 31

**Last Day to add this class this semester:** Jan 31

**Change to Pass/No Pass:** The last day to make this change is Feb 13

**Drop Date for a “W”:** The last day to drop this class is Friday March 14. After this date a grade will be assigned.

**Final Exam Date:** **Wednesday,** **May 21** from11:00-12:50 in CCI 204.

**Grading** : There will be 5 lecture exams and the final exam, equally weighted and counting 65% of your grade. Homework will count 10% and your lab work will count 25% (12.5% lab reports and 12.5% lab quizzes)

General Grading break-off : **A** 90-100%, **B** 80-89%, **C** 70-79%, **D** 60-69%, **F** 0-59%

Please be aware of the following rules:

* Tardiness, talking, leaving early, or sleeping during lectures will result in a partial or full absence being recorded. Students need to sign the sign-in sheet within the first 10 minutes of class.
* Fraudulent behavior during exams is graded with a (0) zero.
* Copying of homework, experimental data, and lab reports is considered fraudulent behavior for both the copier and the originator. DO NOT HAND IN IDENTICAL HOMEWORK.
* No homework may be handed in after I have gone over it in class.  No alternative homework will be given. I will drop the lowest two homework assignments though, lowest 4 if your attendance is good.
* No extra credit will be given. You need to work consistently from the beginning.
* Please turn your cell phones onto “silent buzzer” mode during lectures so as not to disturb the class. No cell phones or i-pods will be allowed during exams.

**LABS**

* Safety glasses need to be worn whenever somebody near you is conducting an experiment.
* No experiments may be conducted without the instructor or teaching assistant present
* No horseplay or unauthorized experiments. Do not taste any chemical or smell any chemical directly.
* Dangerous behavior in the lab will result in the student being asked to leave the lab.
* No visitors inside the lab. You need to go outside to meet with them.
* No food or drinks allowed.
* Backpacks should not be left on the floor where others can trip over them.
* Closed shoes and lab coats must be worn in the lab at all times.
* Long hair should be tied back so it will not fall into chemicals or flames.
* If any accident occurs in the lab, inform your instructor and follow safety procedures. (To be discussed during first lab period)
* Clean up any spills promptly (Clean-up procedures will be discussed during first lab period)
* Do not point the open end of a test tube towards anybody
* Turn off flames when working with organic solvents. Dispose of them in waste bottles in the fume hood, not down the sink.
* At the beginning of each lab your instructor will inform you of any special safety precautions and how to dispose of used chemicals. You need to be on time for the lab so that you hear these instructions.
* Do not dispose of matches, paper or solid chemicals in the sink. Use the large evaporating dishes or sand bucket for spent matches.
* Put broken glassware in the “broken glassware box”, not in the trash.
* Before leaving the lab, wipe the desktop and wash your hands with soap and water.

**If you have a verified need for an academic accommodation (especially in labs) or materials in alternate media (i.e., Braille, large print, electronic text, etc.) per the Americans with Disabilities Act (ADA) or Section 504 of the Rehabilitation Act, please contact the Disabled Student Services as soon as possible.**

**Course Outline: Each Topic takes 1-2 weeks References are to “Chemistry: A Molecular Approach” by Nivaldo Tro.**

A.  Matter and energy (Chapter 1)

      1.   The Laws of conservation of matter and energy (Chapter 1.2)

      2.   States of Matter (Chapter 1.3)

      3.   Chemical and physical properties of matter (Chapter 1.4)

      4.   Chemical and physical changes of matter (Chapter 1.4)

B.   Measurements in chemistry

      1.   Length, mass, volume (Chapter 1.6)

      2.   Density and specific gravity (Chapter 1.6)

      3.   Significant Figures (Chapter 1.7)

      4.   Dimensional Analysis  (Chapter 1.8)

C.   Atoms, molecules, ions, compounds, elements and mixtures (Chapter 2)

      Atomic mass units and isotopes (Chapter 2.6)

D. Nomenclature

      1. Naming inorganic compounds with monatomic and polyatomic ions (Chapter 3.5)

      2. Naming moleculular compounds (Chapter 3.6)

E.   Stiochiometry, chemical formulas, and equations

      1.   Formulas of compounds, etc., and what they mean (Chapter 3.5-3.6)

      2.   The mole, Avogadro’s Number, and molar mass (Chapter 3.7)

      3.   Formula weight, molecular weights, and moles (Chapter 3.7)

      4.   Writing and balancing chemical equations (Chapter 3.10)

      5.   Percent composition and formulas of compound (Chapter 3.8)

             a.    Empirical formula (Chapter 3.9)

             b.   Molecular formula (Chapter 3.9)

                   1)  Chemical equations and calculations (Stoichiometry) (Chapter 4.2)

                   2)  Percent purity, yield, and limiting reagent in equations (Chapter 4.3)

F.   Concentration of solutions (Chapter 4.4)

      1.   Percent by mass and volume

      2.   Molarity (M) molar concentration

      3.   Dilution of solutions

G.  A systematic study of chemical reactions

      1.   Aqueous solutions, electrolytes, nonelectrolytes and extent of ionization (Chapter 4.5)

      2.   Solubility rules (chapter 4.5)

      4.   Net Ionic equations (Chapter 4.7)

      5.   Classification of chemical reactions

             a.    combination and decomposition

             b.   single replacement reactions

             c.    metathesis or double replacement reactions (precipitation, acid-base neutralization) (Chapter 4.6 and 4.8)

             d.   combustion reactions

H.  Acids, Bases, and Salts (Chapter 4.7-4.8)

      1.   Arrhenius acids and bases

      2.   Bronsted-Lowry acids and bases

      3.   Properties of acids and bases

      4.   Preparation of acids and bases

      5.   Concentrations and acid-base reactions in aqueous solutions.

      6.    Titrations

I.    Oxidation Reduction Reaction (Chapter 4.9)

      1.   Assigning oxidation numbers

      2.   Recognizing redox equations by changing in oxidation state

      3.   Balancing simple redox equations (Chapter 18.2)

J.    Physical behavior of gases (Chapter 5)

      1.   The relationship of pressure and volume; Boyle’s Laws (Chapter 5.3)

      2.   The relationship of volume and temperature.  Charles’ Gay Lussac Law (Chapter 5.3)

      3.   Temperature (Kelvin absolute scale) (Chapter 5.3)

      4.   STP : standard temperature and pressure (Chapter 5.3)

      5.   Combined gas laws and molar volume (Chapter 5.3)

      6.    The Ideal Gas Law (Chapter 5.4)

      7.   Molecular weight calculation and Dalton’s Law of partial pressures (Chapter 5.5 and 5.6)

      8.   Graham’s Law of effusion (Chapter 5.9)

K.  Thermochemistry (Chapter 6)

      1.   Heats of reactions and calorimetry (Chapter 6.5-6.6)

      2.   Work (Chapter 6.3)

      3.   The first Law of Thermodynamics (Chapter 6.2)

      4.   Hess’s Law (Chapter 6.8)

      5.   Standard enthalpies of formation (Chapter 6.7-6.8)

L.   Atomic Structure (Chapters 2, 7 and 8)

      1.   Fundamental particles of atom (Chapter 2.6)

      2.   History of atomic structure and fundamental particles  (Chapter 2.4-2.5)

      3.   Atomic number and mass number (Chapter 2.6)

      4.   Nuclear stability and binding energy (Chapter 2.6)

      5.   Atomic spectra and the Bohr atoms (Chapter 7.3)

      6.   Quantum numbers, orbitals, main shells and subshells (7.5-7.6, 8.3)

      7.    Electronic configuration (Chapter 8.3)

M.  Chemical periodicity and ionic bonding

      1.   The periodic table (Chapter 8.2)

      2.   Periodic properties and trends (Chapter 8.6)

      3.   Ionization energy, electron affinity, electronegativity, and size of atoms (Chapter 8.7-8.8)

      4.   Metals, non-metals and metalloids (Chapter 2.7)

      5.    Valence Electrons (Chapter 8.4)

N.  Chemical Bonding (Chapter 9)

      1.   Kinds of chemical bonds

      2.   Ionic bonding, ionic changes, oxidation numbers

      3.   The covalent bond

             a.    polar and nonpolar bonds (Chapter 10.5) and intermolecular forces (Chapter 11.2)

             b.   Lewis dot formulas (Chapter 9.7)

             c.    Octet rule and its limitations (Chapter 9.8? and 9.9)

             d.   Basic motions of bonding theory and resonance (Chapter 9.8)

             e.    Formal charges of Lewis dot formulas (Chapter 9.8)

O.  Covalent bonding and molecular structure (Chapter 10)

      1.   VSEPR Theory and Valence Bond theory (Chapter 10.2-10.3)

      2.   Geometry of molecules from VSEPR or Valence Bond theory (Chapter 10.4)

      3.   Geometry of polyatomic ions (Chapter 10.5)

      4.   The shape of molecular orbitals (Chapter 10.7)

      5.   Energy level diagram of orbitals

      6.   Homonuclear and heteronuclear diatomic molecules (Chapter 10.8)

P.   Liquids and Solids (Chapter 11)

      1.   Liquid state, adhesive and cohesive forces (Chapter 11.3)

             a.    Viscosity

             b.   Surface tension

             c.    Vapor pressure

             d.   Boiliing points and freezing points

             e.    Heat transfer

      2.   The Solid State (Chapter 11.12)

             a.    Melting point

             b.   Heating point

             c.    Sublimation and vapor pressure

             d.   Crystal structure and amorphous

             e.    Bonding in solids

             f.    Metallic bonding (Chapter 9.11)

Q.  Solutions (Chapter 12)

      1.   Solutions terminology

      2.   Concentration units (mole fraction, molality, molarity)

      3.   Dilution of solutions

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| **STUDENT LEARNING OUTCOMES:** |
| ***(Specify the learning skills the student demonstrates through completing the course and link critical thinking skills to specific course content and objectives.)***  |
| Upon completion of this course, students will be able to:  |
| 1. Collect and analyze data and have reasonable conclusions. Assessed by the lab practical.
2. Competent knowledge of the periodic table, molecules, and compounds. Assessed from a pre-test administered at the beginning of the semester and the final exam administered at the end of the semester.
3. Ability to apply skills to solve chemical problems especially math skills. Assessed from a pre-test administered at the beginning of the semester and the final exam administered at the end of the semester.
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| **Chemistry 1A Spring 2014 Cornel** |
| **Week** | **Date** | **Labs (T/Th)** | **Lectures (M/W/F)** |
| 1.Jan 13-17 | Jan 14 | *Lab 1: Introduction to Laboratory Safety* | Syllabus and Periodic Table1. Matter1. Dimensional Analysis  |
| Jan 16 | *Inventory check-in* |
| 2. | Jan 20 | **Martin Luther King Day** | **No lecture Monday**1. Scientific Notation and Significant Figures 2. Atoms |
| Jan 21 | Lab 2: Properties and Changes of Matter |
| Jan 23 | Lab 3: Measurement |
| 3 | Jan 28 | Lab 7: The Mole | 2.9 Mole 3.5 Ionic Compounds 3.6 Molecules and Polyatomic Ions  |
| **Jan 30** | **Lab Quiz 1 (Labs 2, 3, equipment and safety)** |
| 4  | Feb 4 | *Nomenclature Worksheet* | 3.6 More Polyatomic ions 3.10, 4.6 Writing and Balancing Reactions 3.8-9 Empirical Formulas  |
| **Feb 6** | **Thursday Exam 1** |
| 5 | Feb 11 | Lab 6: Empirical Formulas: Oxide of Tin | 4.2 Stoichiometry and Hydrates4.3 Limiting Reactions**No lecture Friday** |
| Feb 13 | Lab 8: The Formula of a Hydrate |
| **Feb 14** | **Lincoln Day** |
| 6 | **Feb 17** | **Washington Day** | **No lecture Monday**4.4 Solutions 4.5 Electrolytes and Net Ionic |
| Feb 18 | Lab 9: Stoichiometry |
| Feb 20 | Lab 10: Alum Crystallization.  Recycling Aluminum Cans |
| 7 | Feb 25 | Lab 5: Double Displacement Reactions  | 4.8 Acid-Base reactions 4.7 Titrations, Reaction Types4.9 Redox Reactions  |
| Feb 27 | Lab 16: Reactions of Copper and finish lab 10 |
| 8 | **Mar 4** | **Tuesday Exam 2**  | 18.2 Balancing Redox 18.2 Redox titrations, Activity series 5. Gas 1 |
| **Mar 6** | **Lab Quiz 2 (Labs 5, 6, 7, 8, 9, 10)** |
| 9  | Mar 11 | Lab 15: Redox Reactions- The Burning of Magnesium and finish lab 11. | 5. Gas 2-4 |
| Mar 13 | Lab 21: Charles's Law |
| **Mar 14** | **Last Day to drop class to get a “W"** |
| 10 | Mar 18 | Lab 22: Molecular Mass of a Volatile Liquid  | 6. Thermo 1-3  |
| Mar 20 | Lab 23: Atomic Mass of an Unknown Divalent Metal |
| 11 | Mar 25 | Lab 27: Heat Flow, Calorimetry | 7. Light 1 and 28.4 Electron Configuration  |
| Mar 27 | Lab 13: Acids and Bases |
| 12 | **Apr 1** | **Tuesday Exam 3** | 7. Quantum Numbers 8. Periodicity9. Lewis Diagrams |
| **Apr 3** | **Lab Quiz 3 (Labs 15, 16, 21, 22, 27)** |
| 13 | Apr 8 | Lab 19: Vitamin C in Fruit Juices | 10. Geometry 1 and 29.8 Formal Charges and Polar Bonds |
| Apr 10 | **Lab Practical (Titration)** |
|  | **Apr 14-18** | **Spring Break** | **No Classes** |
| 14 | Apr 22 | Lab 28: Molecular Geometry Part 1 | 10.5 Dipoles 10.7 Hybridization 9.8 Resonance |
| Apr 24 | Lab 28: Molecular Geometry Part 2 |
| 15 | Apr 29 | Lab 17: Percent Iron(II) in an Unknown | 11.2 Intermolecular Forces 11.2-11.3 Liquids 11.6-8, 11-12 Solids |
| May 1 | Lab 11: Gravimetric Analysis of Phosphorus in Plant Food |
| 16 | May 6 | Lab 29: Polarity | 12.5-6 Solutions 1 and 215. pH and pOH |
| **May 8** | **Thursday Exam 4** |
| 17 | May 13 | Lab 30: Freezing Point Depression and making "Gold" Pennies | **Review for Final** |
| **May 15** | **Lab Quiz 4 (Labs 17, 23, 28, 29)** |
| **18 Wed** | **May 21** | **Final Exam 11:00-12:50pm in CCI 204** |