

CHEM 3A-53007: Introductory General Chemistry Summer 2023

Lectures **M, T, W, Th 10:20am-12:30pm** in MSE 204

Labs **M, T, W, Th 8-10:10am** in MSE 201

You will need to spend at least **2 more hours a day** reading the textbook, doing homework and studying.

If you do not **know how to use Canvas**, please stay after class and ask me to show you!

Instructor: Veronica Cornel

Contact info: e-mail through Canvas or veronica.cornel@reedleycollege.edu using "Chem3A" in subject line, or voicemail 638-3641 ext 3449

Canvas Website: <https://scccd.instructure.com> (All fill-in notes, labs, grades)

Tutoring: **Free in-person in the tutorial center, and free zoom tutoring available:**
<https://www.reedleycollege.edu/academics/tutoring-services/index.html>

Course Objectives: Chemistry 3A is an elementary survey course in chemistry including lab work. It is designed to give the student a chemistry background for a wide variety of careers including forestry, nutrition, nursing, physical therapy, teaching and other biological and health related fields.

Course Prerequisite: Math 3A/High School Algebra II Students will need to be familiar with basic algebra before taking this course as almost half the course is math problems. Stats does not prepare you for the math in this class.

Course Advisories: ENGL 1A, CHEM10 or high school chemistry

Text and Materials:

1. **Optional** but highly recommended: Nivaldo J. Tro: "Introductory Chemistry" **3rd, 4th, 5th or 6th** Edition. The Mastering Chemistry CD is not needed.
2. **Print the fill-in notes and fill-in labs off Canvas before coming to lecture/lab.** You may print them in the library for a small fee (EOPS students and DSP&S students may print in the EOPS and DSP&S offices). Notes will be posted on Canvas on Fridays for the next week.
3. You will need a calculator with "exp" (or "EE") and "log" keys (\$12 at Walmart), but not a programmable calculator, not a CASIO calculator, nor a STATS calculator. You may borrow a calculator for the semester from the library (or EOPS office).
4. I will lend you a labcoat and safety glasses for the course. These will stay in the lab.

Class Attendance and Participation Policy Requirements: Your participation matters! Students who do not attend the first lecture and lab without contacting the instructor may be dropped to allow waitlisted students to register. Students need to attend and pass the lab portion of the class to pass the class. Students who have not attended class or not completed assignments in the class for 3 days, or a cumulative 25%, may be dropped if they do not contact me. This is to enable you to get a "W" rather than a letter grade. If you wish to withdraw from the class, you are responsible for withdrawing before the drop date. After that date you will be given a letter grade.

Late Work Policy: **Homework assignments** need to be turned in at the **beginning of the next lecture** and can only be one lecture late. You will lose 10% if the homework is turned in late (even 10 minutes late!). If you are absent email me your homework through Canvas. If you email me the homework after the due time, but before the next lecture, 10% will still be

deducted for it being late. Prelabs need to be turned in at the beginning of lab (before I start talking), and should not have anything printed on the back. Late prelabs will also lose 10%. If you are late to lab you will also lose points for the lab. Lab reports and Postlab questions need to be turned in before leaving the lab. If you miss a lab, email or turn in the Prelab and Postlab questions before the next lab for a small partial credit.

Quizzes and Exams Late Work Policy: A make-up quiz/exam will only be given due to extreme circumstances, with verifiable written proof, at the discretion of the instructor. Only one make-up quiz/exam will be allowed per student. If you miss a second exam you will get zero for that exam.

No notes, cell phones, i-pods, smart watches, programmable calculators, restroom breaks or talking will be allowed during exams.

Lecture Notes: The ability to listen carefully and to take good lecture notes in an essential college skill. Students should print out the fill-in notes on Canvas and fill them in during the lecture. (If you need an accommodation to have a note-taker or another accommodation, please contact the DSP&S office). The homework is on the last page of the notes and is due at the beginning of the next lecture. Studies have shown that 90% of the lecture material is retained if you do the homework straight after the lecture video and reading the section in your textbook or the free OER textbook on Canvas. If you wait longer to do your homework you will only retain 35% of the lecture. If you get help with your homework, make sure you redo it yourself or do similar problems from your textbook before the next lecture. Keep your notes, homework and lab assignments in a binder and review them the day before your exam, especially the homework and prelabs.

Homework and Lab Reports: Homework will be assigned every lecture and a lab report for every lab. It is essential to your success in this class that you do all the assigned homework and lab reports, and read the relevant sections in your Textbook or the free OpenStax textbook. All homework will due at the beginning of the next lecture and selected problems graded. This is to ensure that you work consistently and can apply what you learn to problems. Do not just copy somebody else's homework or you will not be able to do the problems for yourself in the exams. If you turn in identical homework/lab reports I will give a zero to both the student who copied and the student who let you copy their work. You can ask another student or tutor to help you start some problems, but you need to work them out for yourself. Even if you get all the problems wrong, you will still get 70% for the assignment for attempting all the problems yourself and showing all your work. You will need to print or write out the homework questions/prelab/postlab questions as well as the answers and show your work so you can study your homework/lab report. You can also do the odd number problems in your textbook for extra practice and check the answers at the back of the textbook (or do the problems in the free OER textbook on Canvas).

Last day to add class or to drop class to avoid a "W": June 23, 2023

Drop Date: The last day to drop this class is Friday, July 7. After this date a grade will be assigned.

Laboratory (25%):	Lab Reports 12.5%, Lab quizzes 12.5%
Lecture Material (75%):	Exams 65%
	Homework 10%

The grading scale to be used is **A** 90-100%, **B** 80-89%, **C** 70-79%, **D** 60-69%, **F** 0-59%

CHEM 3A Summer 2023

	<u>Date</u>	<u>Lecture Topic and Chapters in Tro</u>	<u>Lab</u>
Tues	June 20	Syllabus, Periodic Table 2.3 Significant Figures and Density	Lab Safety Periodic Table
Wed	June 21	2.2 Scientific Notation and Temperature 2.6 Dimensional Analysis & Measurements	Exp 3. Densities of Liquids and Solids
Thur	June 22	3. Matter, Physical and Chemical Changes 4. Atoms, Elements and Ions	Exp 1. Properties and Changes of Matter
Fri	June 23	Last day to add the class, or drop the class to avoid a "W"	
Mon	June 26	5. Chemical Nomenclature: Ionic Compounds 5.11 and 6.1-6.1 The Mole	Exam 1
Tues	June 27	5. Chemical Nomenclature: Molecules 6.7 Percent Composition 6.8 Empirical Formulas	Exp 4: The Mole Nomenclature Worksheet
Wed	June 28	5. Polyatomic ions and hydrates Percent water in Hydrates	Exp 5. Empirical Formulas
Thurs	June 29	7.1-7.4 Balancing Chemical Reactions 7.9-7.10 Types of Reactions	Exam 2
Mon	July 3	8. Stoichiometry	Exp 13: Percent water in hydrates
Tues	July 4	No class - Independence Day	
Wed	July 5	7.7 Net Ionic Equations, electrolytes 8.6-7 Limiting Reactants and % Yield	Exp 7: Reaction Types: Copper chemistry
Thurs	July 6	6. Thermochemistry 6. Calorimetry	Exam 3
Fri	July 7	Last day to drop this class to get a "W"	
Mon	July 10	9 Electronic configuration and Orbital Diagrams	Exp 8: Alum production from scrap aluminum Gold and Silver pennies
Tues	July 11	9.7, 9.9 and 10.2 Periodic Table Trends 10. Bonding and 10.5 Lewis Diagrams	Exp 2: Calorimetry
Wed	July 12	10.7 Geometry 10.8 Electronegativity, Polar Covalent Bonds and Polar Molecules	Lewis Diagrams and Molecular Models
Thur	July 13	11. Gases	Exam 4
Mon	July 17	11. Combined Gas law 11. Ideal Gas Law	Lab 10: Charles Law Gummy Bear Demo
Tues	July 18	11. Gas Stoichiometry and Partial Pressure	Exp 14: Molar mass of a volatile liquid
Wed	July 19	12. Liquids, Solids and Intermolecular Forces 13. Solutions, Dilutions	Exp 9: Production of hydrogen gas
Thurs	July 20	14. Acids and Bases 14.6 Titrations	Exam 5
Mon	July 24	14. pH and pOH (bring your calculator)	pH and pOH lab
Tues	July 25	15. Chemical Equilibrium	Exp 11: Acid base titration
Wed	July 26	17. Radioactivity	Making slime
Thurs	July 27	Exam 6	No lab today

Student Learning Outcomes:

Upon completion of this course, students will be able to:

- A. Use dimensional analysis to solve for an unknown parameter of density, volume, mass, pressure, temperature, molar mass, concentration, or an empirical formula.
- B. Construct and balance a chemical reaction and use the reaction to predict stoichiometric quantities.
- C. Explain concepts from the periodic table and use the periodic table to solve chemical problems.
- D. Describe acid-base reactions and how to calculate pH.
- E. Name and draw Lewis diagrams of inorganic and molecular compounds from the formula and vice versa.
- F. Safely conduct laboratory experiments implementing concepts and principles learned in lecture.

Course Objectives:

Upon completion of this course, students will be able to:

- A. demonstrate an appreciation for the impact of chemistry on modern society and the relationship between chemistry and other disciplines including agriculture, the medical field, and industry;
- B. classify types of matter, recognize physical properties and chemical properties, and a general understanding of the Law of Conservation of Mass and the Law of Conservation of Energy;
- C. perform unit conversions using the correct significant figures; between the English and metric systems, temperatures in different units, density, energy, and with SI units;
- D. use the periodic table to predict physical and chemical properties of elements and calculate molar masses of compounds and molecules;
- E. recognize the electromagnetic spectrum and have a basic understanding of the quantum mechanical model of the atom;
- F. demonstrate the ability to name inorganic compounds given their formulas, and write formulas given names;
- G. distinguish and identify metals, non-metals, metalloids, and the elements of alkali metals, alkaline earth metals, halogens, noble gases, transition metals, and elements of the lanthanide and actinide;
- H. distinguish and identify between different types of intramolecular and intermolecular forces of attraction present in various substances based on chemical formulas and structures;
- I. write Lewis Electron-Dot Formulas and identify the shape using VSEPR method;
- J. write and balance chemical equations, and use these equations along with stoichiometry and the mole concept to convert quantities (e.g. grams or moles) of a given substance into quantities of an unknown substance;
- K. calculate, empirical formulas, and percentage composition given the appropriate data;
- L. distinguish and balance chemical equations of different types of reactions;
- M. perform calculations involving a limiting reagent and determining the percent yield;
- N. predict the physical behavior of gases to pressure, temperature, and volume changes;
- O. solve simple mathematical problems involving formula calculations related to gas laws;

- P. use gas laws and stoichiometry to calculate quantities (e.g. moles, volume, grams) of gas produced or consumed during a reaction;
- Q. calculate various parameters of solutions including molarity, dilution techniques, percentage concentration, and density.
- R. construct heating and cooling curves;
- S. describe state and energy changes accompanying heating and cooling curves;
- T. apply the principles of equilibrium in reversible reactions, saturated solutions, solutions of weak electrolytes and solutions of gases in solving related problems;
- U. use solution properties and stoichiometry to calculate quantities (e.g. moles, volume, grams) produced or consumed during a reaction;
- V. describe colligative properties of solutions (e.g. boiling point elevation, freezing point depression, and osmotic pressure);
- W. define and identify acids and bases and be able to perform math calculations involving the pH;
- X. determine the nature and applications for electron exchange reactions;
- Y. understand the structure of the atomic nucleus;
- Z. understand the fundamental types of nuclear radiation and the effects they have on biological systems
- AA. and demonstrate laboratory skills which include operating an analytical balance; calibrating and/or use fundamental lab equipment such as a thermometer, barometer, buret, pipette; recognizing use and limitations of laboratory glassware; recording and reporting observations; using error analysis techniques to evaluate certainty of data; use safety precautions and general laboratory procedures.

Please be aware of the following rules:

- Tardiness, cell-phone use, leaving early, and sleeping during lecture is considered disruptive behavior and will result in a partial or full absence being recorded. Students will need to sign the sign-in sheet within the first 10 minutes of class.
- Fraudulent behavior during exams is graded with a (0) zero.
- Copying of homework, experimental data, and lab reports is considered fraudulent behavior for both the copier and the originator and points (10-100%) may be deducted from both the copier and the originator. **DO NOT HAND IN IDENTICAL HOMEWORK.**
- No homework may be handed in after I have returned it or gone over it in class. No alternative homework will be given. I will drop the lowest two homework assignments though.
- No extra credit will be given. You need to work consistently from the beginning.
- Please turn your cell phones onto “silent buzzer” mode during lectures so as not to disturb the class. No cell phones, i-phones or i-pods will be allowed during exams.

If you have a verified need for an academic accommodation or materials in alternate media (i.e., Braille, large print, electronic text, etc.) per the Americans with Disabilities Act (ADA) or Section 504 of the Rehabilitation Act, please contact the Disabled Student Services as soon as possible.

With this statement on my course syllabus, I am referring each of my enrolled students in need of academic support to tutorial services. Referral reason: Mastering the content, study skills, and basic skills of this course is aided by the use of trained peer tutors